**BOYLES LAWS HOMEWORK**

1. Nitrous oxide (N2O) is used as an anesthetic. The pressure on 2.5L of N2O changes from 105 kPa to 40.5 kPa. If the temperature remains constant what will the new volume be?
2. A gas of 4.00L and pressure of 205kPa is allowed to expand to a volume of 12.0 L. What is its pressure if the temperature remains constant?
3. Calculate the volume of a gas (in L) at a pressure of 1.00 x 102 kPavif its volume at

1.20 x 102kPa is 1.50 x 102 mL.

1. A gas with volume of 4.0L at 90 kPa expands until the pressure drops to 20kPa. What is its new volume if temperature does not change?

Go to following website.

<http://group.chem.iastate.edu/Greenbowe/sections/projectfolder/flashfiles/gaslaw/boyles_law_graph.html>